**Electron configuration of atoms and ions (Level 3) examiners tips: Read these please!**

• electron configuration at Level 3 requires you to write the s, p, d notation

• what is an orbital? a region in space where an electron is most likely to be found

• remember to fill and empty the 4s orbital before 3d

• be careful with both chromium and copper, their stable electronic configurations are written as...

**Cr**: 1s22s22p63s23p64s13d5

**Cu**: 1s22s22p63s23p64s13d10

• the colour of transition metals ions in solution is related to electrons in unfilled d orbitals becoming

excited after absorbing visible light

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